COMPONENTS:

- (1) Gadolinium fluoride; GdF₃; [13765-26-9]
- (2) Dimethylsulfoxide; C_2H_6OS ; [67-68-5]

ORIGINAL MEASUREMENTS:

Kirmse, E.M.

Wiss. Hefte, Paed. Inst. Koethen 1978, 2, 85-90.

VARIABLES:

PREPARED BY:

Room Temperature

T. Mioduski

EXPERIMENTAL VALUES:

The solubility of GdF_3 in $(CH_3)_2SO$ at room temperature was given as

0.035 mass %

The corresponding molality calculated by the compiler is

 $0.0016 \text{ mol kg}^{-1}$

The solid phase was dried in a desiccator over $P_4 O_{10}$ and the Gd:F ratio found to be almost 1:3.

AUXILIARY INFORMATION

METHOD/APPARATUS/PROCEDURE:

Isothermal method. About 100 mg of GdF₃ was added to 10-20 cm³ of solvent, and the mixture mechanically agitated at room temp for 100 h. 5-10 g of saturated solution were removed by decanting or by centrifuging, and the solution evaporated to dryness. The residue was heated with about 10 cm³ of 10% KOH solution for 1-2 h to obtain solid Gd(OH)₃ and a basic F⁻ solution. The precipitate was washed, dissolved in aq HCl, and Gd determined several times by complexometric titration with potentiometric endpoint detection (1). The fluoride content in the filtrate was determined photometrically using Al-Eriochrome cyanine color lake indicator (2).

The reported solubility is a mean of "numerous parallel determinations," or "at least two parallel determinations."

SOURCE AND PURITY OF MATERIALS:

 ${
m Gd_2O_3}$ (source and purity not specified) was dissolved in HCl and the fluoride precipitated by addition of aq HF. The solid produced was ${
m GdF_3.0.5H_2O}$ and was dehydrated by washing with acetone followed by drying at ${
m 310^{\circ}C}$ for 120 hours.

The solvent was dried and purified by "standard methods."

ESTIMATED ERROR:

Soly: results with relative errors exceeding 50% were rejected.

Temp: nothing specified.

REFERENCES:

- Schilbach, U.; Kirmse, E.M. Z. Chem. 1974, 14, 484.
- Schilbach, U.; Hetze, I.; Kirmse, E.M. Chemia Analityczna 1975, 20, 33.